**Ionic and Covalent Bonding Activity**

Step 1:

With your cut-out of two elements, answer the following questions:

|  |  |  |
| --- | --- | --- |
|  | **Metal** | **Non-Metal** |
| Name of element |  |  |
| Atomic number |  |  |
| Atomic mass |  |  |
| Number of protons |  |  |
| Number of neutrons |  |  |
| Number of electrons |  |  |
| Ion charge(s) |  |  |
| Number of valence electrons |  |  |
| Number of electron shells |  |  |
| Bohr diagram |  |  |
| Lewis dot diagram |  |  |

Step 2:

1. Choose another student to bond with.
2. Identify what type of compound you are making (ionic or covalent)
	1. If ionic, label anion and cation (positive and negative charges)
	2. If covalent, label lone pair(s) and bonding pair(s)
3. Draw the bond using a Lewis dot diagram
4. Write out the chemical formula of the compound you have created
5. Repeat this 7 more times

\*\*You must make at least 3 ionic and 3 covalent compounds\*\*

|  |  |
| --- | --- |
| Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: | Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: |
| Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: | Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: |
| Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: | Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: |
| Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: | Type of Compound:Labeled diagramBond using Lewis dot Chemical Formula: |