**3.2 Polyatomic Ions** (p. 91–92)

Some ions are just single atoms. Others are large molecules with several different atoms covalently bonded together. These larger ions are called **polyatomic ions**.

Eg. OH- (hydroxide), CH3COO- (acetate), NH4+ (ammonium), Cr2O72- (dichromate)

1. **NAMING (\**no* –ide endings!)**

**Metal Non-metal{use the polyatomic ion name}**

***See Table 3.10 ON pg. 92*** *or* ***the back side of your periodic table for names***

**Eg.** Sn(OH)2 =

 NH4NO3 =

 Sr(NO3)2 =

 Ca3(PO4)2 =

1. **WRITING FORMULAS**

We write formulas as before, but you MUST use parentheses if the polyatomic ion has a subscript

**Eg.** sodium hydroxide =

 barium sulfate =

 sodium bicarbonate =

 magnesium bisulfite =

**DO pg. 91 Practice Problems**

**Worksheet “BLM 1-38”**

**Quiz \_\_\_\_\_\_\_\_\_**